

Evidence of hollow golden cages

Satya Bulusu, Xi Li, Lai-Sheng Wang, and Xiao Cheng Zeng

PNAS published online May 19, 2006;
doi:10.1073/pnas.0600637103

This information is current as of March 2007.

Supplementary Material

Supplementary material can be found at:
www.pnas.org/cgi/content/full/0600637103/DC1

This article has been cited by other articles:
www.pnas.org/otherarticles

E-mail Alerts

Receive free email alerts when new articles cite this article - sign up in the box at the top right corner of the article or [click here](#).

Rights & Permissions

To reproduce this article in part (figures, tables) or in entirety, see:
www.pnas.org/misc/rightperm.shtml

Reprints

To order reprints, see:
www.pnas.org/misc/reprints.shtml

Notes:

Evidence of hollow golden cages

Satya Bulusu*, Xi Li^{†‡§}, Lai-Sheng Wang^{†¶}, and Xiao Cheng Zeng^{*¶}

*Department of Chemistry and Center for Materials and Nanoscience, University of Nebraska, Lincoln, NE 68588; [†]Department of Physics, Washington State University, 2710 University Drive, Richland, WA 99354; and [‡]Chemical Sciences Division, Pacific Northwest National Laboratory, MS K8-88, P.O. Box 999, Richland, WA 99352

Edited by Benjamin Widom, Cornell University, Ithaca, NY, and approved April 7, 2006 (received for review January 24, 2006)

The fullerenes are the first “free-standing” elemental hollow cages identified by spectroscopy experiments and synthesized in the bulk. Here, we report experimental and theoretical evidence of hollow cages consisting of pure metal atoms, Au_n^- ($n = 16\text{--}18$); to our knowledge, free-standing metal hollow cages have not been previously detected in the laboratory. These hollow golden cages (“bucky gold”) have an average diameter >5.5 Å, which can easily accommodate one guest atom inside.

anion photoelectron spectroscopy | density functional calculation | hollow gold cages | lowest-energy clusters

The isolation and detection of carbon-free hollow cages have attracted much interest since the discovery (1) and synthesis (2) of the buckyball C_{60} and the higher fullerenes. Although “free-standing” inorganic cages have been synthesized (3), bare elemental metal cages have not been observed in nature or detected in the laboratory. Among metals, gold has some unique properties including the strong relativistic effects and aurophilic attraction (4). Recently, a fullerene-like hollow cage with 32 Au atoms was predicted to be highly stable (5, 6). However, photoelectron spectroscopy (PES) combined with theoretical calculations shows that at the relatively large size the overwhelming population of low-lying clusters for Au_{32} near room temperature appears to consist of only compact structures because of the entropic factor (7). Other, larger gold clusters with cage-like local minimum structures also have been suggested (8, 9), but none has been observed experimentally. Conversely, it has been established from both ion-mobility (10) and PES (11) experiments that the most stable anion gold clusters (Au_n^-) in the size range $n = 5\text{--}13$ possess planar structures and that a structural transition from planar to three-dimensional (3D) structures occurs at $n = 14$. Beyond $n = 14$, previous global-minimum searches based on empirical potential functions of gold (12, 13) or semiempirical tight-binding models of gold (14) suggest that all low-lying isomers of gold clusters assume space-filling compact structures. Among the larger gold clusters, Au_{20} is the most interesting; it has been found to possess a pyramidal structure with tetrahedral symmetry just as carved out of the bulk face-centered cubic crystal (15).

Results and Discussion

To elucidate the structural transition from the planar Au at $n = 13$ to the pyramidal Au_{20} , we carried out a joint experimental PES and theoretical study on Au_n^- for $n = 15\text{--}19$. The measured spectra (see *Methods* below) are shown in Fig. 1*A* with numerous well resolved features in the lower binding energy part, which are used to compare with theoretically simulated spectra (Fig. 1*B* and *C* and *Methods* below; see also Fig. 3, which is published as supporting information on the PNAS web site) with the candidate lowest-energy clusters (see Fig. 4, which is published as supporting information on the PNAS web site). The vertical detachment energies (VDEs) (given by the location of the first major peak near the threshold) for this feature are given in Table 1, compared with the theoretical VDEs from the lowest-energy structures. Note that the threshold of the lowest-binding-energy feature in each spectrum (see Table 2, which is published as

supporting information on the PNAS web site) defines the electron affinity of the neutral clusters.

The theoretically obtained top-10 lowest-energy structures (see *Methods*) are given in Fig. 4. Among these top-10 isomers, we selected those isomers within 0.2 eV ($1 \text{ eV} = 1.602 \times 10^{-19} \text{ J}$) from the lowest-energy isomer and simulated their photoelectron spectra (Figs. 1*B* and *C* and 3). We regard these selected isomers as the candidates for the lowest-energy structure owing to the intrinsic error bar ($<0.2 \text{ eV}$) of density-functional theory (DFT) electronic energy calculations (16–18) and the basis-set effects. The number of candidate lowest-energy isomers ranges from one for Au_{19}^- (Fig. 4*E*) to five for Au_{15}^- (Fig. 4*A*) and Au_{16}^- (Fig. 4*B*), and six for Au_{17}^- (Fig. 4*C*) and Au_{18}^- (Fig. 4*D*).

Remarkably, we observed that all but a total of three candidate lowest-energy isomers of Au_{16}^- , Au_{17}^- , and Au_{18}^- are “hollow cages” with an empty interior space (Fig. 4*B–D*). The interior space (typically with length scale >5.5 Å) of these hollow cages can easily host a foreign atom. Among the five candidate lowest-energy structures of Au_{15}^- (Fig. 4*A*), Au_{15a}^- , Au_{15b}^- , and Au_{15d}^- are flat-cage structures, whereas Au_{15c}^- and Au_{15e}^- are pyramid-like structures. Previous studies have shown that in stable gold clusters, gold atoms tend to have maximum coordination number of six, e.g., in the 2D planar structures of $\text{Au}_9\text{--Au}_{13}$ (10, 11) and in the pyramidal structure of Au_{20} (15). Hence, it is understandable that both the flat-cage and pyramid-like structures are energetically competitive for the gold clusters within the size range Au_{14} to Au_{20} . Conversely, it is quite surprising that the hollow-cage structures dominate the low-lying population of Au_{16} to Au_{18} clusters. Specifically, at Au_{16} , only Au_{16c}^- (among the five candidate lowest-energy structures) has flat-cage structure whose interior length scale can be <5 Å (Fig. 4*B*). The isomer Au_{16a}^- can be viewed as a relaxed structure of the pyramidal Au_{20} with four missing corner atoms but maintains the tetrahedral symmetry of Au_{20} (15). At Au_{17} , only Au_{17c}^- among the six candidate lowest-energy structures has a flat-cage structure (Fig. 4*C*), whereas at Au_{18} , only Au_{18a}^- among the six candidate lowest-energy structures exhibits pyramid-like (non-hollow-cage) structure (Fig. 4*D*). Note that Au_{17a}^- can be viewed as placing one atom on top of Au_{16a}^- , whereas Au_{18b}^- can be viewed as placing one atom on top of Au_{17a}^- . Both Au_{17a}^- and Au_{18b}^- possess C_{2v} symmetry. At Au_{19} , there is only a single candidate for the lowest-energy structure, namely, Au_{19a}^- , whose energy is 0.2–0.3 eV lower (depending on the basis set) than the second-lowest-energy isomer (Au_{19b}^-) and $\approx 0.5 \text{ eV}$ lower than the third-lowest-energy isomer (Au_{19c}^-). Au_{19a}^- exhibits a pyramidal structure, which is similar to the pyramidal Au_{20} (15) with one missing corner atom. This structural similarity is expected because Au_{19} is only one atom less than the highly stable (magic-number) pyramidal cluster Au_{20} (15). Compared with Au_{19a}^- , the

Conflict of interest statement: No conflicts declared.

This paper was submitted directly (Track II) to the PNAS office.

Abbreviations: DFT, density-functional theory; PES, photoelectron spectroscopy; VDE, vertical detachment energy.

[§]Present address: Rowland Institute at Harvard, 100 Edwin H. Land Boulevard, Cambridge, MA 02142.

[¶]To whom correspondence may be addressed. E-mail: ls.wang@pnl.gov or xc.zeng@phase2.unl.edu.

© 2006 by The National Academy of Sciences of the USA

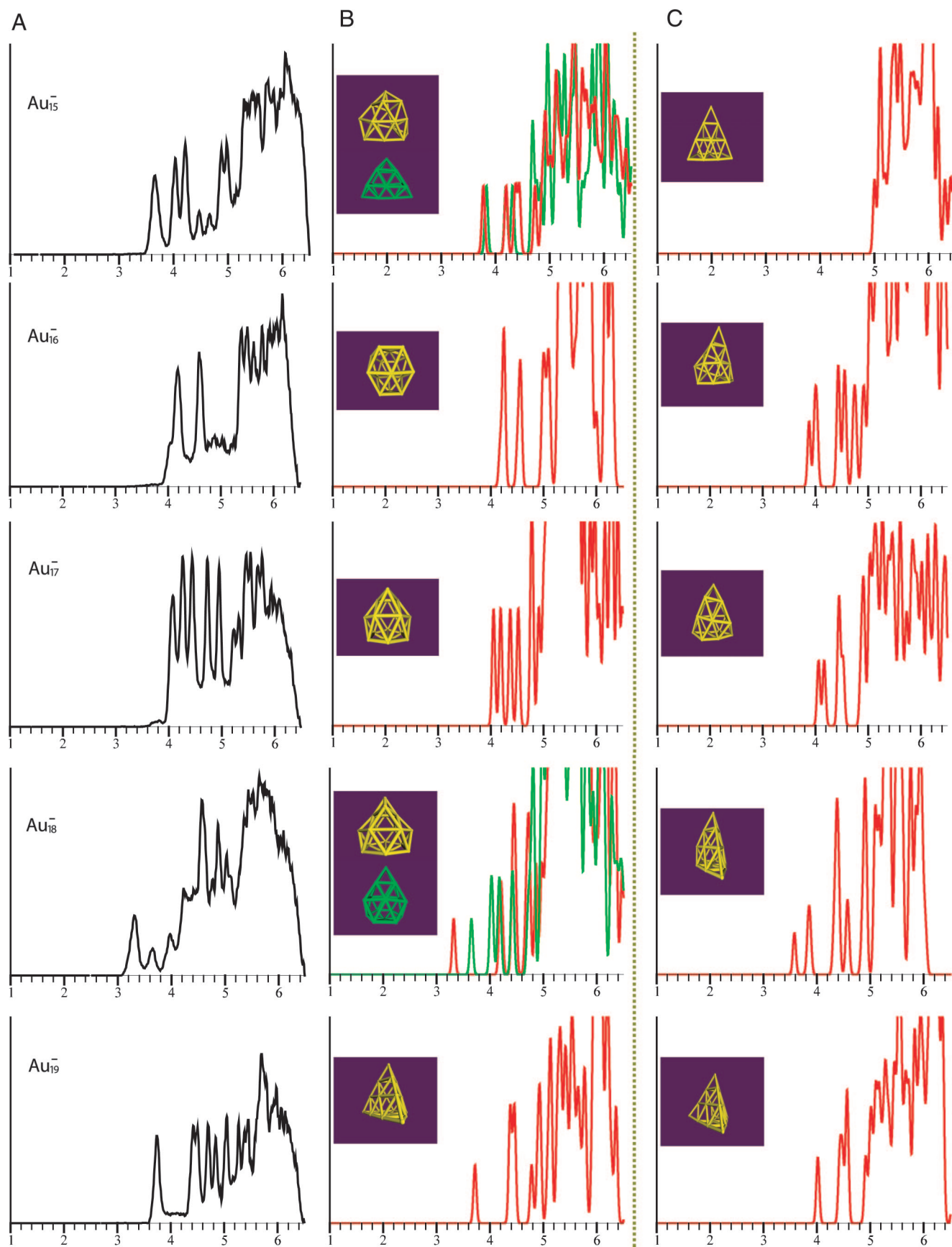


Fig. 1. Experimental photoelectron spectra of Au_n^- ($n = 15\text{--}19$) compared with those simulated theoretically. (A) Experimental spectra measured at 193 nm (6.424 eV). (B) The simulated spectra for one (or two) lowest-lying isomer that matches the first and second major peaks of the measured spectra. (C) The simulated spectra for the non-hollow-cage candidate isomer, which appears not to match the observed spectra.

Table 1. Experimental first VDEs for Au_n^- ($n = 15\text{--}19$) compared with computed values for the candidate lowest-energy clusters that give the best fit to the first two major peaks of the measured spectra

Isomer	VDE	
	Experimental	Theoretical
Au_{15a}^-	3.65 ± 0.03	3.777
Au_{16a}^-	4.03 ± 0.03	4.179
Au_{17a}^-	4.08 ± 0.03	4.053
Au_{18b}^-	3.32 ± 0.03	3.320
Au_{19a}^-	3.74 ± 0.03	3.720

All energies are in eV.

hollow-cage structures such as Au_{19c}^- and Au_{19d}^- are no longer energetically competitive (Fig. 4D). In other words, the structural transition from hollow-cage to pyramid-like structure appears to occur at Au_{19} . To illustrate the structural evolution of gold clusters from 2D planar to 3D flat-cage, hollow-cage, and pyramid-like structures, we highlight in Fig. 2 those candidate lowest-energy clusters that can provide reasonable match to the first two to four major peaks of the experimental photoelectron spectra (Fig. 1A and B).

Our first-principles global search provides the electronic energy-based evidence that the overwhelming majority of the low-lying clusters of Au_{16}^- to Au_{18}^- exhibit hollow-cage structures. Moreover, our measured/simulated PES provides additional spectroscopic evidence to the existence of free-standing hollow golden cages. Here, we used the time-independent DFT (see *Methods*) to obtain approximated theoretical PES for all of the candidate lowest-energy structures of Au_{15}^- to Au_{19}^- (Figs. 1 and 3). Note also that the combined experimental and theoretical PES study has been used by many researchers to explore structures of small- to medium-sized clusters. This approach is particularly effective to identify structures of highly stable (magic-number) clusters such as the buckyball C_{60} or golden pyramid Au_{20} (15) because magic-number clusters are notably lower in energy than other isomers (i.e., they are the undisputed lowest-energy cluster). In this sense, Au_{19a}^- , the sole candidate for the lowest-energy cluster of Au_{19}^- , can be viewed as a

magic-number cluster because of the overwhelming stability of the pyramidal Au_{20} (15). As such, the simulated PES of Au_{19a}^- should match well with the measured one. Indeed, the location of the first two peaks near the threshold, which are directly related to the frontier orbitals and the VDE of the cluster, are in very good agreement with the measured one (including the weak doublet feature of the second major peak). Because the simulated PES based on DFT was obtained from the negatives of the Kohn–Sham (KS) eigenenergies (ground-state energy values), the simulated PES is not expected to match peak-for-peak with the measured PES beyond the threshold (energies of excited states). In summary, the location of the first two major peaks offers a critical structural “fingerprint” of the Au_{19a}^- . Conversely, the simulated PES for the second lowest-energy isomer (Au_{19b}^-), which is also pyramid-like, corresponding to the removal of an atom from the edge of the tetrahedral Au_{20}^- (15), does not agree with the experiment. The VDE of the first simulated peak is too high compared with the experiment (Fig. 1).

For other Au_n^- clusters ($n = 15\text{--}18$), each has five or six candidate lowest-energy structures (Fig. 4). Moreover, previous PES studies of the endohedral gold-cage cluster W@Au_{12} (19, 20) have shown that the gold cage is fluxional. In other words, the energy barriers separating structurally similar isomers (e.g., hollow cages) can be quite small. As a result, it is conceivable that multiple isomers may contribute to the experimental spectra. Hence, our first priority was to use the measured PES as a “filter” to identify those candidate isomers that cannot match the measured PES well (see Fig. 1C). Again, our main focus has been placed on the location of the first two major peaks and, to a lesser extent, the number of peaks in the 4- to 5-eV binding energy range. For example, at Au_{15} , the two pyramid-like low-lying isomers Au_{15c}^- and Au_{15e}^- can be ruled out (Fig. 3A). In fact, the simulated PES of the two flat-cage isomers Au_{15a}^- and Au_{15d}^- seem to match the measured PES (Fig. 1A and B), particularly on the location of the two major peaks near the threshold.

At Au_{16} , the only non-hollow-cage isomer Au_{16c}^- and the isomer Au_{16e}^- can be ruled out because their first VDE seems to be lower than the experimental data (Fig. 3B). The remaining three isomers, Au_{16a}^- , Au_{16b}^- , and Au_{16d}^- , all give reasonable VDE, but Au_{16a}^- seems to provide the best agreement with the experiment in term of the

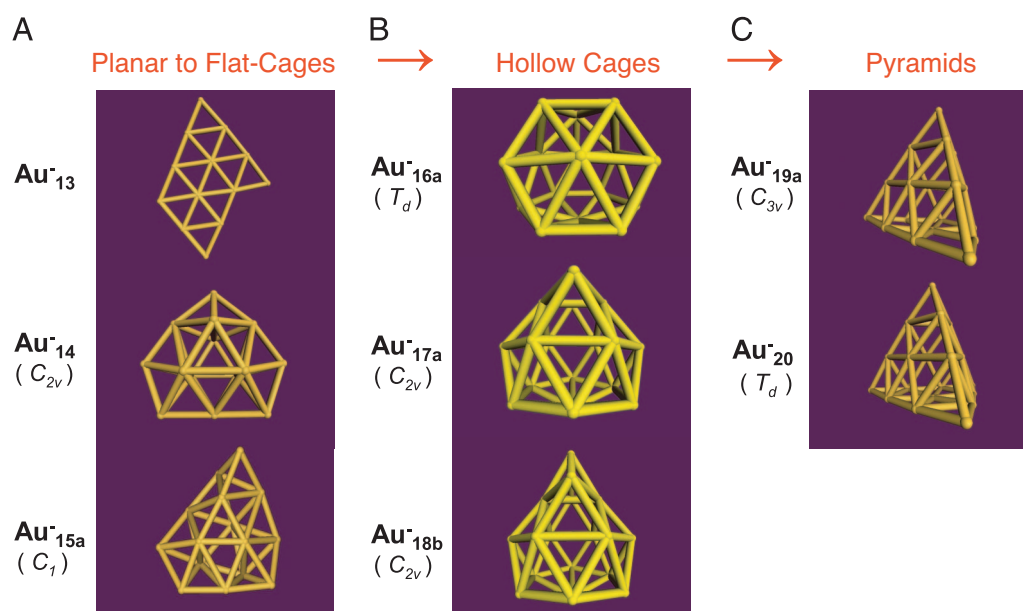


Fig. 2. Structural evolution of mid-sized gold anion clusters from Au_{13}^- to Au_{20}^- . (A) The 2D planar to 3D flat-cage structural transitions (11). (B) The hollow gold cages with diameters >5.5 Å. (C) The pyramid-like clusters, which resemble bulk gold (15).

1. Kroto, H. W., Heath, J. R., O'Brien, S. C., Curl, R. F. & Smalley, R. E. (1985) *Nature* **318**, 162–163.
2. Kratschmer, W., Lamb, L. D., Fostiropoulos, K. & Huffman, D. R. (1990) *Nature* **347**, 354–358.
3. Bai, J., Virovets, A. V. & Scheer, M. (2003) *Science* **300**, 781–783.
4. Pyykkö, P. (2004) *Angew. Chem. Int. Ed.* **43**, 4412–4456.
5. Johansson, M. P., Sundholm, D. & Vaara, J. (2004) *Angew. Chem. Int. Ed.* **43**, 2678–2681.
6. Gu, X., Ji, M., Wei, S. H. & Gong, X. G. (2004) *Phys. Rev. B Condens. Matter* **70**, 205401-1–205401-5.
7. Ji, M., Gu, X., Li, X., Gong, X., Li, J. & Wang, L. S. (2005) *Angew. Chem. Int. Ed.* **44**, 7119–7123.
8. Gao, Y. & Zeng, X. C. (2005) *J. Am. Chem. Soc.* **127**, 3698–3699.
9. Wang, J., Jellinek, J., Zhao, J., Chen, Z., King, R. B. & Schleyer, P. V. R. (2005) *J. Phys. Chem. A* **109**, 9265–9269.
10. Furche F., Ahlrichs R., Weis, P., Jacob C., Glib, S., Bierweiler, T. & Kappes, M. M. (2002) *J. Chem. Phys.* **117**, 6982–6990.
11. Häkkinen H., Yoon, B., Landman, U., Li, X., Zhai, H. J. & Wang, L. S. (2003) *J. Phys. Chem. A* **107**, 6168–6175.
12. Doye, J. P. K. & Wales, D. J. (1998) *New J. Chem.* **22**, 733–744.
13. Wilson, N. T. & Johnston, R. L. (2000) *Eur. Phys. J. D* **12**, 161–169.
14. Wang, J., Wang, G. & Zhao, J. (2002) *Phys. Rev. B Condens. Matter* **66**, 035418-1–035418-6.
15. Li, J., Li, X., Zhai, H. J. & Wang, L. S. (2003) *Science* **299**, 864–867.
16. Olson, R.M., Varganov, R., Gordon, M. S., Metiu, H., Chretien, S., Piecuch, P., Kowalski, K., Kucharski, S. A. & Musial, M. (2005) *J. Am. Chem. Soc.* **127**, 1049–1052.
17. Han, Y.-K. (2006) *J. Chem. Phys.* **124**, 024316-1–024316-3.
18. Yoo, S. & Zeng, X. C. (2005) *J. Chem. Phys.* **123**, 164303-1–164303-6.
19. Pyykkö, P. & Runeberg, N. (2002) *Angew. Chem. Int. Ed.* **41**, 2174–2176.
20. Li, X., Kiran, B., Li, J., Zhai, H. J. & Wang, L. S. (2002) *Angew. Chem. Int. Ed.* **41**, 4786–4789.
21. Yang, M., Jackson, K.A., Koehler, C., Frauenheim, T. & Jellinek, J. (2006) *J. Chem. Phys.* **124**, 024308-1–024308-6.
22. Fernandez, E. M., Soler, J. M., Garzon, I. L. & Balbas, L. C. (2004) *Phys. Rev. B Condens. Matter* **70**, 165403-1–165403-14.
23. Zhai, H. J., Li, J. & Wang, L. S. (2004) *J. Chem. Phys.* **121**, 8369–8374.
24. Gao, Y., Bulusu, S. & Zeng, X. C. (2005) *J. Am. Chem. Soc.* **127**, 15680–15681.
25. Curl, R. F. & Smalley, R. E. (1988) *Science* **242**, 1017–1022.
26. Cai, Y., Guo, T., Jin, C., Haufler, R. E., Chibante, L. P. F., Fure, J., Wang, L., Alford, J. M. & Smalley, R. E. (1991) *J. Phys. Chem.* **95**, 7564–7568.
27. Wang, L. S., Cheng, H. S. & Fan, J. (1995) *J. Chem. Phys.* **102**, 9480–9493.
28. Yoo, S. & Zeng, X. C. (2005) *Angew. Chem. Int. Ed.* **44**, 1491–1495.
29. Perdew, J. P., Burke, K. & Ernzerhof, M. (1996) *Phys. Rev. Lett.* **77**, 3865–3868.
30. Delley, B. (1990) *J. Chem. Phys.* **92**, 508–517.
31. Frisch, M. J., Trucks, G. W., Schlegel, H. B., Scuseria, G. E., Robb, M. A., Cheeseman, J. R., Montgomery, J. A., Jr., Vreven, T., Kudin, K. N., Burant, J. C., et al. (2003) GAUSSIAN 03 (Gaussian, Pittsburgh), Revision C. 02.
32. Dolg, M., Wedig, U., Stoll, H. & Preuss, H. (1987) *J. Chem. Phys.* **86**, 866–872.
33. Schwerdtfeger, P., Dolg, M., Schwarz, W. H. E., Bowmaker, G. A. & Boyd, P. D. W. (1989) *J. Chem. Phys.* **91**, 1762–1774.